ELECTRODE WHICH HAS BEEN COATED WITH A SOLID ION CONDUCTOR WHICH HAS A GARNET-LIKE CRYSTAL STRUCTURE AND HAS THE STOICHIOMETRIC COMPOSITION L_{1-x}A_xG_{3-y}Zr_{y}O_{12}

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References Cited
U.S. PATENT DOCUMENTS
(Continued)

FOREIGN PATENT DOCUMENTS
CN 1694285 A 11/2005
(Continued)

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ABSTRACT
The invention is directed to an article which contains a solid ion conductor having a garnet-like crystal structure and has the stoichiometric composition L_{1-x}A_xG_{3-y}Zr_{y}O_{12}, where L is in each case independently a monovalent cation, A is in each case independently a divalent cation, G is in each case independently a trivalent cation, 0 ≤ x ≤ 3 and 0 ≤ y ≤ 3, wherein the article is a battery, an accumulator, a supercap, a fuel cell, sensor, a thermoelectric converter or an electrochromic device.

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(56) References Cited

U.S. PATENT DOCUMENTS


FOREIGN PATENT DOCUMENTS

JP H1-025956 A 1/1999
JP H1-054147 A 2/1999
JP 2008-050214 1/2008

* cited by examiner
Figure 1

Thick pellet

Experimental
Simulation

18 °C
(5 Hz-13 MHz)

Thin pellet

18 °C

910 kHz
3.3 kHz
Figure 2b

- thin, volume (Ea=0.32 eV)
- thin, total (Ea=0.30 eV)
- thick, volume (Ea=0.34 eV)
- thick, total (Ea=0.31 eV)

log₁₀(σT/Scm⁻¹K⁻¹)

1000T⁻¹/K⁻¹
Figure 6

Thin pellet

(5 Hz-13 MHz)

- 25 °C
- 50 °C
Figure 7

Thin pellet

$log_{10} \sigma T / Scm^{-1}K^{-1}$

1000$T^{-1}/K^{-1}$
Figure 8

a)  
b)  
c)
1 ELECTRODE WHICH HAS BEEN COATED WITH A SOLID ION CONDUCTOR WHICH HAS A GARNET-LIKE CRYSTAL
STRUCTURE AND HAS THE STOICHIOMETRIC COMPOSITION
La₇₋ₓ₉ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋ₓ₆₋
The invention therefore provides, in one embodiment, a solid ion conductor which has a garnet-like crystal structure and has the stoichiometric composition $La_5A_3Ga_3Zr_2O_{12}$ where:

- $L$ is in each case independently a monovalent cation,
- $A$ is in each case independently a divalent cation,
- $G$ is in each case independently a trivalent cation,
- $O_{3x+3}$ and
- $O_{3x+2}$ can be partly or completely replaced by divalent or trivalent anions such as $N^{2-}$.

$L$ is particularly preferably an alkali metal ion, for example $Li^+$, $Na^+$ or $K^+$. In particular, combinations of various alkali metal ions are also possible for $L$. In a particularly preferred embodiment of the invention, $L=Na^+$. Sodium is very inexpensive and available in any amounts. The small $Na^+$ ion can move readily in the garnet-like structures and in combination with zirconium gives chemically stable crystal structures.

$A$ is any divalent cation or any combination of such cations. Divalent metal cations can preferably be used for $A$. Particular preference is given to alkaline earth metal ions such as Ca, Sr, Ba and/or Mg and also divalent transition metal cations such as Zn. It has been found that these ions move very little if at all in the garnet-like compounds according to the invention, so that ion conduction occurs essentially via $L$.

In the above composition, preference is also given to $O_{3x+2}$ and particularly preferably $O_{3x+1}$. In an embodiment according to the invention, $x=0$, so that $A$ is not present in the garnet-like compound.

$G$ is any trivalent cation or any combination of such cations. Trivalent metal cations can preferably be used for $G$. Particular preference is given to $G=La$.

In a structure of the above composition, $O^{2-}$ can be partly or completely replaced by other anions. For example, it is advantageous to replace $O^{2-}$ completely or partly by other divalent anions. Furthermore, $O^{2-}$ can also be diolavently replaced by trivalent anions with appropriate charge compensation.

In a further aspect, the present invention provides a solid ion conductor of the stoichiometric composition $La_5A_3La_{3-x}Zr_2O_{12}$, where $A$ is a divalent metal and $L$ is $Li$ or $Na$. Because of its ready availability, $Na$ is particularly preferred. In a preferred embodiment, $x=0$, so that the composition is $La_5La_3Zr_2O_{12}$.

$A$ is preferably selected from among alkaline earth metals, preferably from among Ca, Sr, Ba and/or Mg. Preference is likewise given to $A$ being selected from among divalent transition metals, for example $A=Zn$. Greatest preference is given to $A=Sr$ or $Ba$.

Ion conductors of the composition $La_5A_3La_{3-x}Zr_2O_{12}$ have a garnet-like crystal structure. Compared to the known compounds of the composition $La_5LaNb_2O_{12}$ ($L=Li$), the two $Nb^{4+}$ cations have formally been replaced by two $Zr^{4+}$ cations and two monovalent $L$ cations. In addition, $La^{3+}$ may have been replaced by $A^{2+}$ and $L^{+}$. In this way, the total proportion of $L$ in the structure has been increased. $L$ is preferably Li or Na, via which the ion conduction of the compounds having a garnet-like structure occurs. As a result, the compounds of the present invention make it possible to provide significantly improved ion conductors.

Compared to the compounds of the prior art, the materials of the composition $La_5A_3La_{3-x}Zr_2O_{12}$ display an increased ion conductivity. Owing to the garnet structure of the compounds of the present invention, which is a 3D-isotropic structure, ion conduction in three dimensions without a preferential direction is possible.

The electronic conductivity of the compounds of the present invention is, on the other hand, comparativel low. The polycrystalline samples of the compounds of the present invention also have a low grain boundary resistance, so that the total conductivity is made up virtually exclusively of the volume conductivity.

A further advantage of the materials is their high chemical stability. The materials display, in particular, no discernible changes on heating in contact with molten lithium. At temperatures up to 350°C and DC voltages up to 6 V, no chemical decomposition is observed.

An example of a particularly preferred compound according to the invention having a garnet structure is $Li_4La_3Zr_2O_{12}$. The high lithium ion conductivity, good thermal and chemical stability in respect of reactions with possible electrodes, environmental compatibility, availability of the starting materials, low manufacturing costs and simple production and sealing make $Li_4La_3Zr_2O_{12}$ a promising solid electrolyte which is particularly suitable for rechargeable lithium ion batteries.

According to a further aspect, the present invention provides a process for preparing the solid ion conductors having a garnet-like structure. The compounds can be formed by reaction of appropriate salts and/or oxides of the elements present, for example by means of a solid-state reaction. Particularly useful starting materials are nitrates, carbonates and hydroxides which are converted into the corresponding oxides during the course of the reaction.

The present invention more specifically relates to a process for preparing the solid ion conductors of the composition $La_5A_3Ga_{3-x}Zr_2O_{12}$ (e.g. $Na_9Al_{2}Zr_2O_{12}$, $Al_2Zr_2O_{12}$). The materials can be obtained by reaction of appropriate nitrates and/or oxides of $A$, $G$ and $Zr$, with oxides of $Na$ and $Zr$ with a hydroxide, nitrate or carbonate of $L$ in a solid-state reaction. $A$ is as defined above. The divalent metal $A$ is preferably used in the form of nitrates. Here, preference is given to $Ca(NO_3)_2$, $Sr(NO_3)_2$ and $Ba(NO_3)_2$. In the case of $G$, preference is given to using $La$ which is preferably employed in the form of $La_2O_3$, $Zr$ is advantageously used as oxide, preferably $ZrO_2$. $L$ is preferably used in the form of $LiOH$, $LiNO_3$ or $Li_2CO_3$. For example, $LiOH$, $H_2O$ or $NaOH$ can preferably be used.

To compensate for a weight loss of $L$ (e.g. $L=Li$, $Na$) during the heat treatment of the samples, the respective salt is preferably used in excess, for example an excess of 10% by weight.

The starting materials are mixed in a first step and can, for example, be milled in 2-propanol in a ball mill using zirconium oxide milling media. The mixture obtained in this way is subsequently heated at temperatures in the range of preferably 400-1000°C in air for a number of hours, preferably 2-10 hours. Temperatures of 600-800°C, for example about 700°C, and a heat treatment time of 4-8 hours e.g. about 6 hours, are particularly suitable. Milling is then carried out again, preferably likewise in 2-propanol in a ball mill using zirconium oxide milling media. The reaction product is subsequently pressed uniaxially or preferably isostatically to give moulded pieces, for example pellets. These are then sintered for a number of hours, preferably 10-50 hours, more preferably 20-30 hours, at temperatures in the range of preferably 700-1200°C, more preferably 800-1000°C. Temperatures of about 900°C and a heat treatment time of about 24 hours are particularly suitable here. During this sintering process, it is advantageous to cover the samples with a powder of the same composition in order to avoid excessive losses of the $L$ oxide.
Possible methods which can easily be employed for preparing the compound are precursor methods, e.g. the Pecchina method, the glycine method or precipitation reactions, since soluble salts exist for all components.

The solid ion conductors of the invention (e.g. lithium or sodium ion conductors) are, as solid-state electrolytes, a valuable starting material. Since the materials have an extraordinarily high ion conductivity accompanied by negligible electron conduction, they can be used as solid electrolyte for batteries (e.g. lithium or sodium batteries) having a very high energy density. The high stability of the materials in respect of chemical reactions, e.g. with elemental lithium and customary electrode materials, leads to, for example, the solid ion conductors of the present invention being able to be put to practical use in batteries.

The resistance of the phase boundary between the solid electrolytes of the present invention and the electrodes is also very small compared to customary solid electrolyte materials. As a result, batteries having a comparatively high power (high currents) can be produced using the materials according to the invention. The use of the solid-state electrolytes of the present invention also results in improved safety compared to the use of liquid electrolytes. This is of particular advantage when the electrolytes are used in motor vehicles.

In a further aspect, the present invention provides for, apart from the use in batteries, the use of the solid ion conductors (e.g. lithium ion conductors) in electrochromic systems (windows, VDUs, exterior walls, etc.) and for instantaneous energy storage and release in supercapacitors (supercaps). When the ion conductors of the invention are used, it is possible to achieve energy densities of capacitors of 100 F/cm² or more. A further aspect of the invention is the use of the garnet-like solid ion conductors as sensors, in particular for numerous gases. According to the invention, it is also possible to use the material in thermoelectric converters for efficient direct conversion of heat into electric energy.

The ion conductors having garnet-like structures can also be used as buffer layers in combination with other electrolytes, for example conventional aprotic liquid electrolytes. It is therefore not necessary to use an electrolyte which consists entirely of the garnet-like structure. Rather, it is possible to use any known electrolytes which can, for example, be present in liquid, gel or solid form in combination with the novel garnet-like ion conductors.

The invention therefore provides, in a further aspect, for the use of a solid ion conductor having a garnet-like crystal structure as protective layer before an electrode so as to improve the chemical stability towards the electrolyte. For this purpose, it is possible to use not only the garnet-like structures according to the invention containing zirconium but also, for example, the garnet-like compounds described in WO 2005/085138. The use of the ion conductors as buffer structure before the electrodes prevents short circuits and makes it possible to generate and apply relatively high voltages so as to achieve a significantly greater energy density and life.

**A BRIEF DESCRIPTION OF THE FIGURES**

**FIG. 1:**
AC impedance curve of Li₃La₃ZrO₁₂, measured at 18°C in air on a thick pellet (1.02 cm thick and 0.92 cm in diameter). The continuous line represents the simulated data for an equivalent current circuit comprising (R₂Q₂) (R₂₀Q₂₀) (Q₂₀) (where R is the resistance and Q is the constant phase element and the indices g, gb and el indicate grain volume, grain boundary and electrode) using the EQUIVALENT Program (B. A. Bouchamp, Equivalent Circuit, Version 4.05, 1997, Faculty of Chemical Technology, University of Twente, 7500 AE Enschede (The Netherlands), Report No. CT88/265/128/CT89/214/128, May 1989). The impedance curve measured at 18°C in air on a thick pellet (0.18 cm thick and 0.98 cm in diameter) of Li₃La₃ZrO₁₂ is shown in the inset.

**FIG. 2:**
a) Arrhenius curves for the electrical volume and total conductivity (volume and grain boundaries) of the thick pellets of Li₃La₃ZrO₁₂, obtained in two successive heating and cooling cycles.
b) Comparison of the Arrhenius curves obtained for the thick and thin pellets of Li₃La₃ZrO₁₂ during the first heating run (18-300°C).

**FIG. 3:**
Comparison of the total conductivity (volume+grain boundaries) of Li₃La₃ZrO₁₂ and other known lithium ion conductors which come into question for battery applications.

**FIG. 4:**
Measured powder XRD pattern of Li₃La₃ZrO₁₂ and standard pattern of the known garnet phase Li₃La₃Nb₂O₁₂ (JCPDS: 80-0457) as per Joint Committee on Powder Diffraction Standards.

**FIG. 5:**
AC impedance curves measured at 25 and 50°C in air on the thick pellet of Li₃La₃ZrO₁₂.

**FIG. 6:**
AC impedance curves measured at 25 and 50°C in air on the thin pellet of Li₃La₃ZrO₁₂. The further curve at higher frequency is shown as an inset.

**FIG. 7:**
Arrhenius curves for the electrical volume and total (volume+grain boundary) conductivity of the thick pellet of Li₃La₃ZrO₁₂, obtained in two successive heating and cooling cycles.

**FIG. 8:**
Photographs of a) Li₃La₃ZrO₁₂ pellet and molybdenum crucible before exposure to molten lithium, b) Li₃La₃ZrO₁₂ pellet in molten lithium and c) Li₃La₃ZrO₁₂ pellet and molybdenum crucible immediately after exposure to molten lithium for 48 hours. The photographs depicted in Figure c) shows that the colour of the pellet remains unchanged (ivory colour) and no reaction product is formed.

**DETAILED DESCRIPTION OF THE INVENTION**

The following example serves to illustrate a particularly preferred embodiment of the present invention.

**EXAMPLE**

Stoichiometric amounts of the in each case highly pure starting materials:
LiOH (Alfa Aesar, >99%), predried at 200°C for 6 h, 10% by weight excess in order to compensate for the Li loss during the sintering process; 
La₂O₃ (Alfa Aesar, >99.99%), predried at 900°C for 24 h; and 
ZrO₂ (Alfa Aesar, >99%) were reacted in a solid-state reaction.

The starting materials were ball-milled for about 12 hours in 2-propanol using zirconium oxide containers and balls.

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This was followed by heat treatment at 900 and 1125°C in air for 12 hours. The product obtained was then once again ball-milled. The reaction products were subsequently iso-
statically pressed to form pellets and sintered at 1230°C for 36 h. The samples were covered with a powder having the same composition during this procedure in order to avoid an excessive loss of lithium. The heating rate in all treatments was 1°C per minute. The sintered compressed pellets were cut into thinner pellets by means of a diamond saw. Phase formation was monitored using X-ray powder diffraction (XRD) (SEIFERT 3000, CuKα, Germany). The lattice constants were determined from the powder XRD data using the method of least squares.

The measurement of the electrical conductivity was car-
ried out in air using two pellets of differing thickness (thick pellet: 1.02 cm thick and 0.92 cm in diameter, and thin pellet: 0.18 cm thick and 0.98 cm in diameter). The mea-
surement was carried out using Li-ion-blocking Au elec-
trodes (Au paste cured at 700°C for 1 h) in the temperature range from 18 to 350°C by means of an impedance and gain phase analyzer (HP 4192 A, Hewlett-Packard Co., Palo Alto, Calif.) (5 Hz–13 MHz). Before each impedance measurement, the samples were equilibrated at constant temperature for from 3 to 6 hours. The impedance measurements were carried out in two successive heating and cooling cycles for each pellet. The data for the thermogravimetric analysis (TGA) and differential thermal analysis (NETZSCH STA 449C/18) were measured in air at 10°C/min at 350°C.

The stability of Li₂La₂Zr₂O₁₂ towards molten lithium was examined in an argon-filled glove box by reacting the pellet with a large excess of molten lithium in a molybdenum crucible for 48 hours.

Although numerous X-ray diffraction (XRD) studies have been carried out on Li₂La₂M₂O₇ (M=Na, K) garnets, there has been controversy about the structure in respect of the space group and position of the lithium cations (a) D. Mazza, Mater. Lett. 1988, 7, 205-207; b) H. Hiyama, K. Hayashi, Mater. Res. Bull. 1988, 23, 1399-1407; c) J. Isuzu, M. L. Veiga, R. Sanchez-Puche, A. Jerebic, C. Pico, J. Alloys Compd. 1991, 177, 251-267. Recently, neutron diffraction studies have indicated that Li₂La₂M₂O₇ (M=Na, K) crystallizes in the space group P3m1 and that Li is located both in the tetrahedral positions and octahedral positions and that vacancies are present in both types of positions (a) E. J. Cussen, Chem. Commun. 2004, 412-413; b) M. P. O'Callaghan, D. R. Lynham, E. J. Cussen, G. Z. Chen, Chem. Mater. 2006, 18, 4681-4689.

The measured powder XRD pattern of Li₂La₂Zr₂O₁₂ agrees well with the standard pattern of the known garnet phase Li₂La₂M₂O₇ and demonstrates the ability of the garnet structure to incorporate cations of differing oxidation state and different size without an excessive change in the symmetry. The diffraction pattern for a cubic cell having a lattice constant of a = 12.9682 Å was determined.

A typical impedance curve obtained at 18°C for a thick pellet of Li₂La₂Zr₂O₁₂ is shown in Fig. 1. The occurrence of the rise in the region of low frequencies when the electrodes are ionically blocked indicates that the material examined is an ion conductor (a) V. Thangadurai, R. A. Huggins, W. Weepner. J. Power Sources 2002, 108, 64-69; b) J. T. S. Irvine, D. C. Sinclair, A. R. West, Adv. Mater. 1990, 2, 132-138). Similar behaviour has been observed for the previously studied materials having a garnet-like structure. The impedance curve could be resolved into volume, grain boundary and electrode resis-
tances. The continuous line in Fig. 1 represents the data for an equivalent circuit current of (R₁Q₁)(R₂Q₂)(R₃Q₃) using the EQUIVALENT Program. The impedance curve for the thin pellet of Li₂La₂Zr₂O₁₂ measured at 18°C is shown as an inset in Fig. 1. The volume and total conductivity of the thick pellet (1.02 cm thick and 0.92 cm in diameter) and the thin pellet (0.18 cm thick and 0.98 cm in diameter) of Li₂La₂Zr₂O₁₂, observed at various temperatures were obtained from the intersections of the high-frequency and low-frequency semicircles with the axes and are summarized in Table 1. The data shown in Fig. 1 and Table 1 indicate similar electrical properties for the thick and thin pellets of Li₂La₂Zr₂O₁₂. The thin pellet displays a slightly higher volume and total conductivity compared to the thick pellet. In addition, it is interesting to note that the grain boundary contribution to the total resistance is less than 5% and decreases with increasing temperature (Table 1) both for the thick pellet and for the thin pellet. At higher temperature (above 75°C for the thick pellet and above 50°C for the thin pellet), it is difficult to determine the grain boundary contribution compared to the volume contribution precisely; the total value of the volume and grain boundary contributions has therefore been shown for the determination of the electrical conductivity over the temperature range examined. The total conductivity at room temperature (3x10⁻⁷ S/cm at 25°C) of the novel crystalline fast lithium ion conductor Li₂La₂Zr₂O₁₂ having a garnet-like structure is better than that of all other solid lithium ion conductors and all previously described lithium garnets.

This result, viz. that total and volume conductivities are of the same order of magnitude, is a particularly advantageous property of the Li₂La₂Zr₂O₁₂ garnet structure examined here as compared to other ceramic lithium ion conductors. For many applications of solid electrolytes in electrochemical devices such as batteries, sensors and electrochromic displays, the total conductivity should be as high as possible. In addition, volume and total conductivity can be improved further by means of a low-temperature synthesis of Li₂La₂Zr₂O₁₂ and by means of further densification using a suitable sintering process. The Arrhenius curves for the electrical volume and total conductivity of Li₂La₂Zr₂O₁₂, obtained in two heating and cooling cycles, are shown for the thick pellet in Fig. 2a. There is no appreciable change in the conductivity between the two cycles. This means that the garnet-like structure examined is thermally stable and that no phase transition occurs in the temperature range examined, viz. from room temperature to 350°C. A similar Arrhenius behaviour was also observed for the thin pellet of Li₂La₂Zr₂O₁₂. In Fig. 2b, the data for the thick pellet and the thin pellet of Li₂La₂Zr₂O₁₂, which were in each case obtained in the first heating run, are compared. The activation energies obtained for the volume and total conductivity of the thick pellet (0.32 eV at 18-50°C and 0.30 eV at 18-300°C) are slightly lower than those for the volume and total conductivity of the thick pellet (0.34 eV at 18-70°C and 0.31 eV at 18-300°C). The conductivity obtained for the thin pellet is slightly higher than that of the thick pellet.

In addition to the impedance analysis, the ionic nature of the electrical conduction was also confirmed by the measurements in which Li₂La₂Zr₂O₁₂ was used as solid electrolyte between elemental lithium and Al. LiAl. The sample was covered with an aluminium layer on the upper side and placed on lithium which had been melted in a glove box filled with inert Ar gas. The aluminium was alloyed both by chemical reaction with lithium and also by coulometric titration of lithium into the aluminium from the lithium.
e electrode located opposite it. The resulting voltage was in the region of the theoretical value. The difference could be attributed to the inhomogeneous temperature distribution and corresponding phenomena due to irreversible processes.

FIG. 3 shows a comparison of the lithium ion conductivity of Li$_3$La$_3$Zr$_2$O$_{12}$ with other known lithium ion conductors which are under consideration for use in connection with batteries. The conductivity is higher than that of Li$_6$-alumina, thin-layer LiPon (Li$_2$PO$_3$N$_{0.65}$), Li$_3$SiAlO$_9$, Li$_4$40 mol Al$_2$O$_3$, LiZr$_2$(PO$_4$)$_3$, Li$_3$Sb$_2$O$_7$, PO$_4$, Li$_3$La$_3$Ta$_2$O$_{12}$ and Li$_3$Ba$_3$Ta$_2$O$_{12}$. The high lithium conductivity compared to other lithium-containing garnets and low activation energy which are observed are possibly due to the increase in the cubic lattice constants, the increase in the lithium ion concentration, a reduced chemical interaction between the lithium ions and other ions which form the lattice and partly to the improved densification (92% of the theoretical density). At relatively low temperatures, the conductivity of the less stable polycrystalline Li$_3$N (6.6x10$^{-4}$ S/cm at 725°C) is comparable with that of Li$_3$La$_3$Zr$_2$O$_{12}$. However, at higher temperature, Li$_3$La$_3$Zr$_2$O$_{12}$ displays a higher total conductivity.

The fundamental stability of Li$_3$La$_3$Zr$_2$O$_{12}$, which is the thermal stability of the Li$_3$N crystalline lithium ion conductor, was confirmed by thermogravimetric measurements (TGA) and differential thermal analysis (DTA). The TGA-DTA data measured in an air atmosphere indicated no significant change in the mass and no discernible phase change both during heating and during cooling with the temperature range from 20 to 900°C. It was found that zincium-containing Li$_3$La$_3$Zr$_2$O$_{12}$ is stable towards molten lithium and is also chemically stable to the action of moisture and air over the observation period of several weeks.

TABLE 1

<table>
<thead>
<tr>
<th>Type of pellet</th>
<th>Temperature (°C)</th>
<th>$\sigma_{\text{elect}}$ (Scm$^{-1}$)</th>
<th>$\sigma_{\text{total}}$ (Scm$^{-1}$)</th>
<th>$R_{\text{en}}/R_0 + R_{\text{gb}}$</th>
<th>$R_{\text{en}}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>Thick</td>
<td>18</td>
<td>3.73 x 10$^{-4}$</td>
<td>1.90 x 10$^{-4}$</td>
<td>0.44</td>
<td>0.44</td>
</tr>
<tr>
<td>pellet</td>
<td>25</td>
<td>4.67 x 10$^{-4}$</td>
<td>2.44 x 10$^{-4}$</td>
<td>0.48</td>
<td>0.48</td>
</tr>
<tr>
<td>Thin</td>
<td>18</td>
<td>3.97 x 10$^{-4}$</td>
<td>2.32 x 10$^{-4}$</td>
<td>0.42</td>
<td>0.42</td>
</tr>
<tr>
<td>pellet</td>
<td>25</td>
<td>5.11 x 10$^{-4}$</td>
<td>3.01 x 10$^{-4}$</td>
<td>0.44</td>
<td>0.44</td>
</tr>
</tbody>
</table>

1) $R_{\text{en}}$ = grain boundary resistance, $R_0$ = volume resistance

TABLE 2

<table>
<thead>
<tr>
<th>h</th>
<th>k</th>
<th>l</th>
<th>d$_{hkl}$ (Å)</th>
<th>d$_{lkl}$ (Å)</th>
<th>l$_{hkl}$</th>
</tr>
</thead>
<tbody>
<tr>
<td>2</td>
<td>1</td>
<td>1</td>
<td>5.278</td>
<td>5.294</td>
<td>99</td>
</tr>
<tr>
<td>2</td>
<td>2</td>
<td>0</td>
<td>4.574</td>
<td>4.584</td>
<td>17</td>
</tr>
</tbody>
</table>

The invention is claimed:
1. An article which comprises a solid ion conductor which has a garnet-like crystal structure and has the stoichiometric composition Li$_3$La$_3$Zr$_2$O$_{12}$, wherein L is in each case independently a monovalent cation, A is in each case independently a divalent cation, G is in each case independently a trivalent cation, $0{\alpha}_{\times}{\alpha}$3 and O can be partly or completely replaced by divalent or trivalent anion, wherein the article is a battery, an accumulator, a supercap, a fuel cell, sensor, a thermoelectric converter or an electrochemical device.
2. The article according to claim 1, wherein the article is a battery.
3. The article according to claim 1, wherein L is Li$^+$, Na$^+$ or K$^+$.
4. The article according to claim 1, wherein A is Sr or Ba, $0{\alpha}_{\times}{\alpha}$1, L is Li$^+$ or Na$^+$ and G is La.
5. The article according to claim 1, wherein G is La.
6. The article according to claim 1, wherein A is Zn, Ca, Sr, Ba or Mg.
7. The article according to claim 1, wherein A is Sr or Ba, $0{\alpha}_{\times}{\alpha}$1, L is Li$^+$ or Na$^+$ and G is La.